



Supplementary Figure 1 | Deformation Mechanism. **a-b**, Motion of atoms in the auxetic and non-auxetic 1T-MX₂ compounds. The solid and open blue arrows indicate the applied force direction and resulting atom motion direction, respectively. The dashed red arrow indicates the direction of the t_{2g} -p orbital coupling force. **c**, Change in $\angle M_1XM_2$ after step-ii relaxation (i.e., the relaxation with the lattice b fixed to its value in the relaxed strain-free structure). **d**, Change in X-M₃-M₁-M₂ dihedral angle $\angle XQM_2 - (\angle XQM_2)^\circ$ after step-iii relaxation. The $(\angle XQM_2)^\circ$ and $(\angle M_1XM_2)^\circ$ are the reference angles taken in the relaxed structure of corresponding strain-free MX₂. The diamond and circle symbols represent the angles extracted from the structures relaxed under fixed lattice- b and unfixed lattice- b , respectively, subjected to a 5% tension strain in the lattice- a direction. The solid lines and arrows in c-d are guides to the eyes.

Supplementary Table 1: The Poisson's ratios and lattice constants of groups 6-7 MX₂ compounds calculated from SCAN and HSE06. Available experimental lattice constants are shown in parentheses.

	In-plane Poisson's ratio along a : SCAN/HSE06			Lattice Constant a (Å): SCAN/HSE06 (Experiment)		
$M^{4+} \backslash X_2$	-S ₂	-Se ₂	-Te ₂	-S ₂	-Se ₂	-Te ₂
Mo	-0.07/ 0.01	-0.12/-0.00	-0.09/-0.01	3.1998/3.1657 (3.201) ¹	3.2685/3.2529 (3.270) ¹	3.4970/3.4626 (3.469) ²
W	-0.04/ 0.05	-0.19/-0.05	-0.20/-0.06	3.1908/3.1843 (3.222) ¹	3.2574/3.2645	3.4970/3.4759 (3.496) ³
Tc	-0.10/ 0.03	-0.04/ 0.01	-0.37/-0.02	3.0692/3.0255	3.1543/3.1386	3.4149/3.3486
Re	-0.13/-0.02	-0.03-0.02	-0.22/-0.09	3.0750/3.0590	3.1311/3.1381	3.3834/3.3551

Supplementary Table 2. SCAN-calculated lattice constants of 42 1T-MX₂ compounds. Available experimental values are shown in parentheses. Note the experimental lattice constants listed for d⁵-d⁶ MTe₂ compounds are extracted from the distorted 1T structure. The small difference in lattice constant between the undistorted and distorted phases is expected.

		Lattice Constant a (Å)		
M^{4+}	X_2	-S ₂	-Se ₂	-Te ₂
d ⁰	Ti	3.4055 (3.405) ⁴	3.5165 (3.535) ⁴	3.7648 (3.766) ⁴
	Zr	3.6833 (3.662) ⁴	3.7815 (3.770) ⁴	4.0064 (3.950) ⁴
	Hf	3.6153 (3.635) ⁴	3.7180 (3.748) ⁴	3.9606 (3.949) ⁵
d ¹	V	3.2668 (3.29) ⁶	3.3260(3.352) ⁴	3.6022 (3.595) ⁴
	Nb	3.3870	3.4845	3.6738 (3.642) ⁷
	Ta	3.3524 (3.346) ⁴	3.4602 (3.477) ⁴	3.6702 (3.651) ⁷
d ²	Mo	3.1998 (3.201) ¹	3.2685 (3.270) ¹	3.4970 (3.469) ²
	W	3.1908 (3.222) ¹	3.2574	3.4970 (3.496) ³
d ³	Tc	3.0692	3.1543	3.4149
	Re	3.0750	3.1311	3.3834
d ⁵	Co	Non-Layered Structure		3.5983 (3.802) ⁸
	Rh			3.7563 (3.92) ⁸
	Ir			3.8431 (3.928) ⁸
d ⁶	Ni	3.3174	3.4712	3.7248 (3.854) ⁸
	Pd	3.5408	3.6759	3.9162 (4.037) ⁸
	Pt	3.5237 (3.543) ⁴	3.6662 (3.728) ⁴	3.9554 (4.026) ⁸

Supplementary references

- 1 Gordon, R., Yang, D., Crozier, E., Jiang, D. & Frindt, R. Structures of exfoliated single layers of WS₂, MoS₂, and MoSe₂ in aqueous suspension. *Physical Review B* **65**, 125407 (2002).
- 2 Dawson, W. & Bullett, D. Electronic structure and crystallography of MoTe₂ and WTe₂. *Journal of Physics C: Solid State Physics* **20**, 6159 (1987).
- 3 Brown, B. E. The crystal structures of WTe₂ and high-temperature MoTe₂. *Acta Crystallographica* **20**, 268-274 (1966).
- 4 Wilson, J. & Yoffe, A. The transition metal dichalcogenides discussion and interpretation of the observed optical, electrical and structural properties. *Advances in Physics* **18**, 193-335 (1969).
- 5 Smeggil, J. & Bartram, S. The preparation and x ray characterization of HfTe_{2-x}, x= 0.061. *Journal of Solid State Chemistry* **5**, 391-394 (1972).
- 6 Gamble, F. Ionicity, atomic radii, and structure in the layered dichalcogenides of group IVb, Vb, and VIb transition metals. *Journal of Solid State Chemistry* **9**, 358-367 (1974).
- 7 Brown, B. E. The crystal structures of NbT₂ and TaTe₂. *Acta Crystallographica* **20**, 264-267 (1966).
- 8 Jobic, S., Brec, R. & Rouxel, J. Anionic polymeric bonds in transition metal ditellurides. *Journal of Solid State Chemistry* **96**, 169-180 (1992).